**SI WORKSHEET 11**

1. (at this stage) What makes something an acid? What makes something a base?
2. What are the two products of an acid-base reaction?
   1. H2SO4 + KOH 🡪
   2. What is the net ionic equation?
3. What is the equivalence point? What is the difference between this and the end point?
4. What volume is needed to make a .75M NaOH if you have 50 grams of it?
5. You have a stock solution of 18 M Sulfuric acid however for the reaction you need to perform at this time, 18M is too concentrated. However, you find out that you can perform the reaction with 100ml of 3M Sulfuric acid. What volumetric amount do you need of 18M Sulfuric acid to make 100ml of 3M?
6. What amount of .1M NaOH would it take to neutralize 100 ml of .25M HCl?
   1. What is the concentration of NaOH if .150L of .3M HCL was neutralized by 100 ml of NaOH?
7. A 95% ethanol solution has a density of .839 g/ml. What is the molarity of the solution? (\*for this remember how you used percents when you did empirical formulas\*)