**SI WORKSHEET 15**

1. C3H8 + O2 🡪 CO2 + H2O ⍙H=? (\*Appendix J)
   1. What is the enthalpy if 100 g of H2O are formed?
   2. How many grams of propane would be required to give off 1000kJ of energy?
2. Target reaction: C4H10 + O2 🡪 CO2 + H2O ⍙H=?
   1. C(s) + H2 🡪 C4H10 ⍙H= -103.8 kJ
   2. H2 + O2 🡪 H2O (g) ⍙H= -241.8 kJ
   3. C + O2 🡪 CO2 ⍙H= -393.5 kJ
3. Sr + C + O2 🡪SrCO3(s) ⍙H=?
   1. Sr + O2 🡪SrO ⍙H= -592 kJ
   2. SrO + CO2 🡪 SrCO3 ⍙H= -234 kJ
   3. C + O2 🡪 CO2 ⍙H= -393.5 kJ
4. If C10H22(l) + O2 🡪 CO2 + H2O(g) ⍙H= -3000 kJ
   1. C(s) + H2 🡪 C10H22 ⍙H= ?
   2. H2 + O2 🡪 H2O (g) ⍙H= -241.8 kJ
   3. C + O2 🡪 CO2 ⍙H= -393.5 kJ
   4. What would be the enthalpy of reaction for .1 moles of decane?
5. OF2 + H2O(g) 🡪HF(g) + O2 ⍙H=?
   1. The standard molar enthalpy of formation of OF2 is 18kJ/mol, calculate the standard molar enthalpy change for the reaction.