**SI WORKSHEET 5**

1. A compound has a molar mass of 46 g/mol and is 34.8% oxygen, 13% hydrogen, and 52.8% carbon, what is this compound?
2. Given: H2  + O2 🡪 H2O. How many moles of water can be produced from 5 grams of Hydrogen gas?

**2-7: Give the type of reaction for each question**

1. AlPO4 + NaCl 🡪

What are the products of this reaction and how would this equation be balanced?

If starting with 100 grams of NaCl, how many grams of the product with Sodium is formed?

1. Balance: C6H12O6 + O2 🡪 H2O+ CO2 and state how many grams of water are produced from 100 g of glucose (C6H12O6)
2. H2SO4 + Cu(OH)2 🡪 H2O + CuSO4.
	1. Given 200 grams of Copper (II) hydroxide, how many molecules of CuSO4 are produced?
3. CaCO3🡪
	1. If 150 grams of the **product with Calcium** was produced how many grams of CaCO3 did you start with?
4. C10H20 + O2 🡪 CO2 + H2O.
	1. If 9.87 x 1021 molecules of Oxygen are present, how many grams of CO2 is produced?
5. H2SO4 + Ca(OH)2 🡪
	1. If starting with 100 grams of both reactants which one produces more of the products (in grams)?