**SI WORKSHEET 6**

1. CaCO3🡪
   1. If 150 grams of the **product with Calcium** was produced how many grams of CaCO3 did you start with?
2. C10H20 + O2 🡪 CO2 + H2O.
   1. If 9.87 x 1021 molecules of Oxygen are present, how many grams of CO2 is produced?
3. C10H18 + O2 🡪
   1. If 50 grams of cyclodecene are reacted with excess oxygen how much of each product is produced in grams? If the reaction had a 75% yield, how many grams of each product was actually produced?
4. Al2(SO4)3 + NaClO4 🡪
   1. If 100 grams of each reactant is present, how much of the excess reagent remains at end of reaction? How much of each product is produced (in grams) if there’s a 70% yield?
5. C6H12O6 + C2H5OH 🡪 C8H16O6 + H2O
   1. If there are 20 ml of ethanol and 50 grams of glucose, how many water molecules can you produce? (density of ethanol =0.789 g/ml)
6. CaC2O4 + Na3PO4 🡪
   1. A student performs this reaction with 75 grams of Calcium oxalate and 125 g of Sodium phosphate and produces 45 g of the calcium-containing product and the student determines the percent yield to be 65%. Is the student correct? If not, what should the student have received for percent yield?