**SI WORKSHEET 8**

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| Ion/ compound | Solubility |
| Group 1a/ NH4+ |  |
| ClO3-/ ClO4- |  |
| NO3- |  |
| CH3COO- |  |
| C2O42- |  |
| CO32- |  |
| PO43- |  |
| S2- |  |
| OH- |  |
| Cl-, Br-, I- |  |

1. Name 7 strong acids and 7 strong bases.
2. State if soluble or not, and the reason why the compound is soluble or insoluble:
	1. Ammonium phosphate
	2. Iron (III) Chloride
3. Potassium Sulfide
4. Barium Sulfate
5. Calcium phosphate
6. AlCl3 + K3PO4 🡪
7. Knowing this reaction has a 70% yield how much precipitant is formed from 100 grams of each reactant?
8. What is the net ionic equation
9. H2SO4 + NaHCO3 🡪
10. What kind of equation is this?
11. Provide the total and net ionic equations
12. FeCl3 + Co3(PO4)2 🡪 Fe3(PO4)2 + CoCl3
	1. What is the net ionic equation?
13. Think about it: CH3COOH + NH3 🡪 CH3COO- + NH4+. Will there be more products or more reactants?